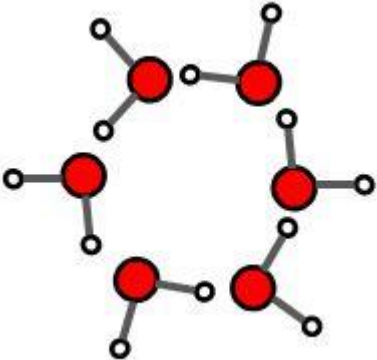
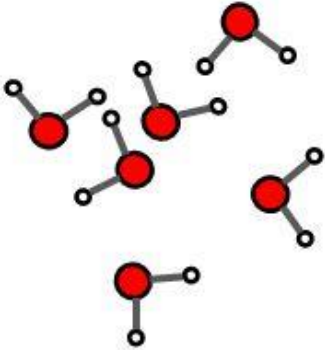

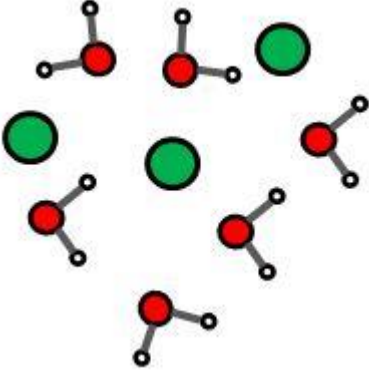


Solid H <sub>2</sub> O (ice)	Liquid H <sub>2</sub> O (water)
	
Solid NaCl solution (salt ice)	Liquid NaCl solution (salt water)
	

- 1. Ice melts when heat speeds up water molecules so fast they cannot align into solid “crystal” structure (a structure like the hexagon above).**

I think the salt makes it *easier* to melt ice.

I think the salt makes it *harder* to melt ice.

- 2. Pure water starts to melt when the temperature reaches 32 °F.**

I think that salt ice starts to melt when it is *colder* than 32 °F.

I think that salt ice starts to melt when it is *warmer* than 32 °F.

- 3. Ice melts when energy from the surroundings is transferred to the water molecules in order to speed them up.**

I think that melting salt ice will *remove* energy from the surroundings, making it *colder*.

I think that melting salt ice will *add* energy to the surroundings, making it *warmer*.

## Predictions

Which ice solution will melt first? Which last?

Which ice solution will make the surrounding melted water coldest? Which warmest?

## Experiment

	pure ice	salt	flour	sugar
melting order				
final temperature (°F)				